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**Synthesis and Characterization of Cyclotriphosphazene containing silicon
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Synthesis and Characterization of Cyclotriphosphazene containing silicon as single solid state precursors for the formation of silicon/phosphorus nanostructured materials

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and Colm O'Dwyer[§]

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Abstract

The synthesis and characterization of new organosilicon derivatives of N₃P₃Cl₆, N₃P₃[NH(CH₂)₃Si(OEt)₃]₆ (**1**), N₃P₃[NH(CH₂)₃Si(OEt)₃]₃[NCH₃(CH₂)₃CN]₃ (**2**) and N₃P₃[NH(CH₂)₃Si(OEt)₃]₃[HOC₆H₄(CH₂)CN]₃ (**3**) is reported. Pyrolysis of (**1**), (**2**) and (**3**) in air and at several temperatures results in nanostructured materials whose composition and morphology depends on the temperature of pyrolysis and the substituents of the phosphazenes ring. The products stem from the reaction of SiO₂ with P₂O₅ leading to either crystalline Si₅(PO₄)₆O, SiP₂O₇ and/or an amorphous phase as the glass Si₅(PO₄)₆O/3SiO₂·2P₂O₅ depending on the temperature and nature of the trimer precursors. From (**1**) at 800 °C, core-shell microspheres of SiO₂ coated with Si₅(PO₄)₆O are obtained while in other cases mesoporous or dense structures are observed. Atomic force microscopy examination after deposition of the materials on monocrystalline silicon wafers evidences morphology strongly dependent on the precursors. Isolated islands of size ~9 nm are observed from (**1**) whereas dense nanostructures with a mean height of 13 nm are formed from (**3**). Brunauer-Emmett-Teller measurements show mesoporous materials with low surface areas. The proposed growth mechanism involves the formation of cross-linking structures and of vacancies by carbonization of the organic matter, where the silicon compounds nucleate. Thus, for the first time, unique silicon nanostructured materials are obtained from cyclic phosphazenes containing silicon.

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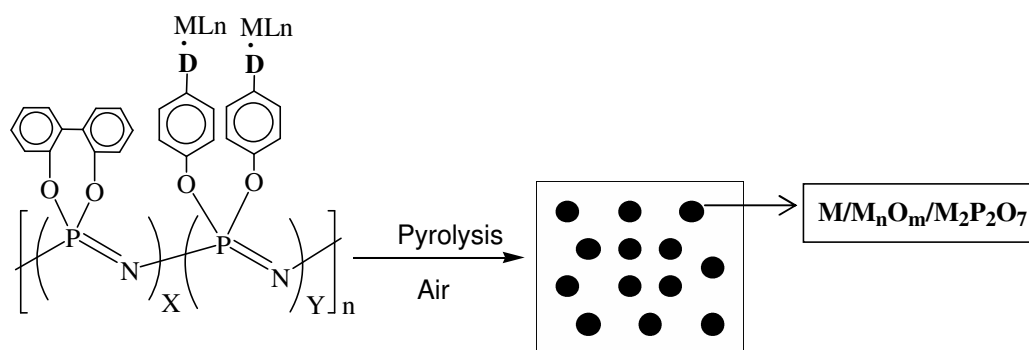
Introduction

Silicon-based compounds are of immense technological importance from monocrystalline silicon as the basis of the digital age to polycrystalline silicon in photovoltaic devices¹. The fabrication and study of silica (SiO₂) has received considerable attention in recent years, largely due to their potential in diverse applications (e.g. catalysis^{2,3}) other than in silicon-based electronic architectures. Their nanoscale counterparts enjoy resurgence in interest due in part to the fundamental differences in properties between the nanoscale and the bulk material⁴.

Nanoparticles of silicon are known to exhibit particle size-dependent optical and electronic properties. Such properties are thought to have important applications in the development of optoelectronic devices and as solubilized crystalline silicon⁵. Also, nanostructured silica has received considerable attention due to exhibited potential in applications such as photonic / optics crystals, nanomicroelectronics/photronics, bionanotechnology and nanocatalysis⁶.

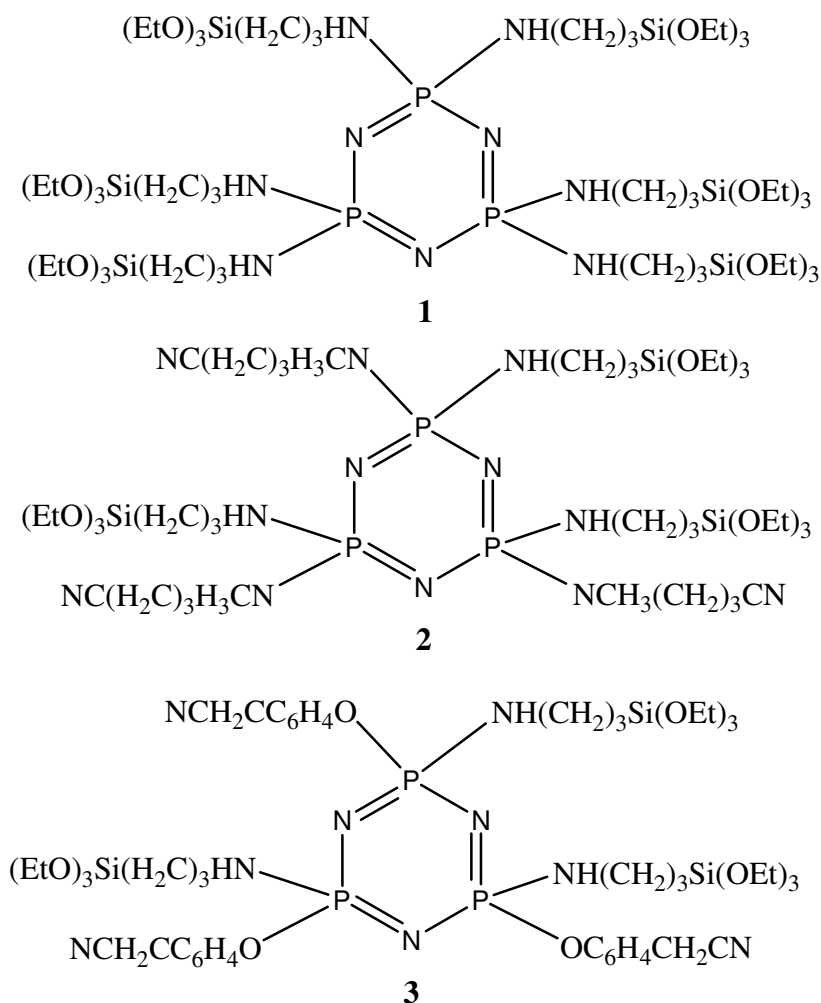
Several preparation methods of silicon nanoparticles have been reported; the majority facilitated by solution reduction of SiCl₄⁷. On other hand, silicon dioxide is generally prepared by the sol-gel method^{8,9}. The majority of methods for obtaining either Si or silica nanoparticles are based in solutions using polymers as a stabilizer. Apart from Si and SiO₂ nanoparticles, few other Si containing nanoparticles have been reported.

We have recently reported¹⁰ a solid-state method of obtaining metallic M⁰; M_xO_y and M_x(P₂O₇)_y nanoparticles by solid-state pyrolysis of polyphosphazene-containing organometallic derivatives, (solid-state pyrolysis of organophosphazene/organometallic, SSPO method), summarized in Scheme 1.



Scheme 1. Schematic representation of the SSPO method.

Because of the importance in obtaining silicon, silica and silicon-containing nanostructured materials in their solid state, we have successfully developed a synthetic method of preparing solid state silicon-containing nanomaterials from pyrolysis of organosilicon derivatives of cyclotriphosphazenes; the formula structure of each organosilicon precursor is outlined in Scheme 2.



Scheme 2 Structural formulae of the new organosilicon cyclotriphosphazenes (1), (2) and (3).

Although polyphosphazenes have been previously used to stabilize gold nanoparticles in solution¹¹, no examples of stabilization of metal nanoparticles by cyclic phosphazenes in solution or in solid-state have been reported. We previously reported on the synthesis and characterization of organometallic derivatives of cyclotriorganophosphazene¹² and here we detail the synthetic method and characterize several new organosilicon derivatives of cyclotriphosphazenes and their pyrolysis

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3 study. Although several cyclotriphosphazenes containing silicon substituents have been prepared by
4 Allcock et al.¹³ to our best knowledge, the synthesis of precursors (1)-(3) has been not reported.

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6 For the first time, nanostructured silicon-containing materials are obtained from solid-state
7 pyrolysis of silicon derivatives of cyclotriphosphazenes.
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10 11 12 **Experimental Section**

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15 All reactions were carried out under dinitrogen using standard Schlenk techniques. Infra-red (IR)
16 spectra were recorded on an FT-IR Perkin-Elmer 2000 spectrophotometer. Solvents were dried and
17 purified using standard procedures using $N_3P_3Cl_6$, $[NBu_4]Br$, K_2CO_3 , $H_2N(CH_2)_3Si(OEt)_3$,
18 $HN(CH_3)(CH_2)_3CN$ and $HOC_6H_4CH_2CN$ (Sigma-Aldrich). Nuclear magnetic resonance (NMR)
19 spectra were conducted using a Bruker AC-300 instrument with $CDCl_3$ as the solvent unless
20 otherwise stated. 1H and $^{13}C\{^1H\}$ NMR are given in δ relative to TMS. $^{31}P\{^1H\}$ are given in δ
21 relative to external 85% aqueous H_3PO_4 . Coupling constants are in Hz. Thermogravimetric analysis
22 (TGA) and differential scanning calorimetry (DSC) measurements were performed on a Mettler TA
23 4000 instrument and Mettler DSC 300 differential scanning calorimeter, respectively. The trimer
24 samples were heated at a rate of $10^\circ C\ min^{-1}$ from ambient temperature to $1000^\circ C$ under a constant
25 flow of nitrogen.
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29 X-ray diffraction (XRD) was carried out at room temperature on a Siemens D-5000
30 diffractometer with θ - 2θ geometry. The XRD data was collected using $Cu-K\alpha$ radiation (40 kV and
31 30 mA). Scanning electron microscopy (SEM) and energy dispersive X-ray analysis (EDX) were
32 acquired with a JEOL 5410 SEM with a NORAN Instrument micro-probe transmission microscope.
33 Transmission electron microscopy (TEM) were carried out on a JEOL SX100 TEM and on a JEOL
34 JEM-2011 operating at 200 kV. The finely powdered samples were dispersed in *n*-hexane and
35 dropped on a conventional carbon-coated copper grid dried under a lamp. The pyrolysis experiments
36 were carried out by pouring a weighed portion (0.05–0.15 g) of the organometallic trimer into
37 aluminum oxide boats placed in a tubular furnace (Lindberg/Blue Oven model STF55346C-1) under
38 a flow of air, heated from 25 to $300^\circ C$ and then to $800^\circ C$, and annealed for 2 h. The heating rate was
39 $10^\circ C\ min^{-1}$ under an air flow of $200\ mL\ min^{-1}$. Brunauer-Emmett-Teller (BET) surface areas were
40 calculated from the adsorption isotherm, using a Micromeritics ASP 2010 instrument. The pore size
41 distribution was evaluated using the BJH method.
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Atomic force microscopy (AFM) measurements were performed using a Veeco Explorer AFM in tapping mode. Roughness, feature size and image treatment was conducted using the accompanying software. Suitable depositions were obtained by dissolving the trimers (1), (2) and (3) in dichloromethane, dropping on the silicon wafer, followed by evaporation of the solvent at room temperature. Subsequent pyrolysis was conducted at 800°C.

Synthesis of $N_3P_3[NH(CH_2)_3Si(OEt)_3]_6$ (1)

A solution of $N_3P_3Cl_6$ (5 g, 14.4 mmol) in toluene (20 ml) was added dropwise to 3-aminopropyl(triethoxy)silane (21.03 g, 95 mmol) and triethylamine (10.57 g, 105 mmol) in toluene (50 ml) at room temperature. The reaction mixture was heated under reflux for 4 h. During this time a white precipitate formed. The precipitate was removed by filtration. The solvent was evaporated from the filtrate and the 3-aminopropyl(triethoxy)silane and triethylamine residues were evaporated under vacuum at 100°C. The product was obtained as an opaque liquid in quantitative yield.

Elemental analysis: Calc for $C_{54}H_{132}N_9O_{18}P_3Si_6$ (found) C: 44.5 (40.00); H: 9.13 (9.60); N: 8.65 (8.26). ^{31}P NMR (ppm, $CDCl_3$) = 0.72. 1H NMR (ppm, $CDCl_3$): δ = 1.18 m, 72 H; $[(CH_2)_3Si(OCH_2CH_3)_3]_6$; 0.77 m, 54 H. $[(CH_2)_3Si(OCH_2CH_3)_3]_6$
IR (KBr, cm^{-1}) 3373 $\nu(N-H)$; 2975, $\nu(C-CH_3)$ 2928, 2886, $\nu(C-CH_2)$; 1483, $\delta(Si-CH_2)$; 1189, 1162, $\nu(PN)$; 1103, 1081, 957 $\nu(Si-O)$, 791, 772 $\delta(C-CH_3)$. Mass spectrum m/z 1455(M^+), 1410($M^+ - OCH_2CH_3$), 1248($M^+ - OCH_2CH_3 - Si(OCH_2CH_3)_3$).

Synthesis of $N_3P_3[NH(CH_2)_3Si(OEt)_3]_3[N(CH_3)(CH_2)_2CN]_3$ (2)

A solution containing 5 g (14.4 mmol) of $N_3P_3Cl_6$, 20.21 g (86.14 mmol) of 3-aminopropyl(triethoxy)silane, 8.066 ml (86.14 mmol) of $HN(CH_3)(CH_2)_2CN$ with triethylamine (14.43 ml, 104.0 mmol) in 40 ml of toluene was heated under reflux for 6 h. The light-brown precipitate that formed was separated by filtration through a neutral alumina column and the solvent removed from solution. The 3-aminopropyl(triethoxy)silane and triethylamine residues were evaporated under vacuum at 100°C. The product was obtained as red-brown oil in quantitative yield.

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3 Elemental analysis: Calc. For $C_{42}H_{93}N_{12}O_9P_3Si_3$ (found) C: 44.6 (42.21); H: 8.88 (9.44); N: 10.62
4 (12.6). ^{31}P NMR (ppm, $CDCl_3$) = 20.8 m 17.97 m ppm. 1H NMR (ppm, $CDCl_3$): δ = 1.20 m,
5 $[(CH_2)_3Si(OCH_2CH_3)_3]_3$ 36 H; 0.58 m, 27 H, $[(CH_2)_3Si(OCH_2CH_3)_3]_3$. ^{29}Si NMR ($CDCl_3$, ppm) =
6 -45.36 (Si(OCH_2CH_3))₃. IR (KBr, cm^{-1}) 3354, 3222, $\nu(N-H)$; 2974, $\nu(C-CH_3)$ 2927, 2885, $\nu(C-CH_2)$;
7 2247.6, $\nu(CN)$ 1443, $\delta(Si-CH_2)$; 1190, 1167, $\nu(PN)$; 1102, 108, 957 $\nu(Si-O)$, 791, 770 $\delta(C-CH_3)$.
8 Mass spectrum m/z 1086 (M^+), 1086 ($M^+ - 3OCH_2CH_3 - Si(OCH_2CH_3)_3$), 989($M^+ - N(CH_3)(CH_2)_3CN$
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21 **Synthesis $N_3P_3 [NH(CH_2)_3Si(OEt)_3][OC_6H_4(CH_2)CN]_3$ (3)**

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24 To a solution of $N_3P_3Cl_6$ (5 g, 14.4 mmol) in toluene (20 ml), 3-aminopropyl(triethoxy)silane
25 (1.58 g, 7.15 mmol) and triethylamine (14.43 g, 104 mmol) in toluene (50 ml) was added dropwise at
26 room temperature. The reaction mixture was heated under reflux for 1.5 h. The white precipitate was
27 filtered off and the solution heated at 100°C to eliminate the 3-aminopropyl(triethoxy)silane and
28 triethylamine residues. Subsequently, a solution of 0.95 ml (7.15 mmol) of $HOC_6H_4CH_2CN$, 11.63 g
29 (35.75 mmol) of potassium carbonate and 11.52 g (35.75 mmol) of $[NBu_4]Br$ was added and the
30 mixture heated under reflux in acetone for 6 h. The resulting brown solid was removed by filtration
31 and the solution was chromatographed on neutral alumina and eluted with CH_2Cl_2 . The solvent was
32 removed from the eluate and the red-brown oil was dried under vacuum for 6 h.
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40 Elemental analysis: Calc. for $C_{51}H_{84}N_9O_{12}P_3Si_3$ (found) C: 46.67 (50.53); H: 8.53 (11.18); N:
41 9.42 (6.82). ^{31}P NMR($CDCl_3$) = 20.89, 18.69. 1H NMR (ppm, $CDCl_3$): δ = 0.97 m,
42 $[(CH_2)_3Si(OCH_2CH_3)_3]_3$ 36 H; 1.19 m; 27 H, $[(CH_2)_3Si(OCH_2CH_3)_3]_3$. IR (KBr, cm^{-1}) 3417, $\nu(N-$
43 H); 2967, $\nu(C-CH_3)$ 2934, 2876, $\nu(C-CH_2)$; 2240, $\nu(CN)$; 1486, $\delta(Si-CH_2)$; 1189, 1167, $\nu(PN)$; 1104,
44 1082, 958, $\nu(Si-O)$; 793, 771 $\delta(C-CH_3)$. Mass spectrum m/z 1191 (M^+), 1059 ($M^+ - OC_6H_4CH_2CN$),
45 844 ($M^+ - OCH_2CH_3 - Si(OCH_2CH_3)_3$)
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55 **Results and discussion**

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Synthesis of the precursors

The analytical data are not perfect for compounds **(1)**, **(2)** and **(3)**. This can be due to two main factors: the known hydrolytic instability of compounds having the $-\text{Si}(\text{OEt})_3$ moiety⁸ as well as the incomplete combustion of these type of samples to give carbon content lower than that calculated. On the other hand attempts to purification of the oil samples by distillation gave rise to decomposition.

Reaction of $\text{N}_3\text{P}_3\text{Cl}_6$ with $\text{H}_2\text{N}(\text{CH}_2)_3\text{Si}(\text{OEt})_3$ in the presence of triethylamine in toluene as a solvent, and at reflux yields $\text{N}_3\text{P}_3[\text{NH}(\text{CH}_2)_3\text{Si}(\text{OEt})_3]_6$ (**1**) as viscous in color liquid. The ^{31}P -NMR spectrum exhibits the typical signal at 0.72 ppm characteristic of $\text{N}_3\text{P}_3(\text{NR}_2)_6$ derivatives¹² (the spectrum is shown in supporting information S₁). The ^1H -NMR spectrum exhibited the expected signal of the aminopropyl(triethoxy)silane¹⁵ group see experimental part. In the IR spectrum the $\nu(\text{PN})$ ring vibration at 1189 and 1162 cm^{-1} ¹², and the $\nu(\text{Si-O})$ at 791, 772 cm^{-1} ^{16c} and the $\nu(\text{NH})$ at 3373 cm^{-1} , were observed. Mass spectrum of the product shows the expected molecular ion, see experimental part. Another fragments arising from loss of OCH_2CH_3 and $\text{Si}(\text{OCH}_2\text{CH}_3)_3$ were also observed.

The reaction of $\text{N}_3\text{P}_3\text{Cl}_6$ with $\text{H}_2\text{N}(\text{CH}_2)_3\text{Si}(\text{OEt})_3$ and $\text{HN}(\text{CH}_3)(\text{CH}_2)_3\text{CN}$ in the presence of triethyl amine and toluene as a solvent results in $\text{N}_3\text{P}_3[\text{NH}(\text{CH}_2)_3\text{Si}(\text{OEt})_3]_3[\text{NCH}_3(\text{CH}_2)_3\text{CN}]_3$ (**2**) as a red viscous liquid. Unlike to compound **(1)**, the ^{31}P NMR spectrum for the compound **(2)** showed two resonances at 20.8 ppm. and 17.97 ppm(both nearly multiplets) with a relative intensity of 1:2, see supporting information , S₂. In the three stages of substitution in $\text{N}_3\text{P}_3\text{Cl}_6$, two products are expected^{17,18}, the geminal and nongeminal compound. For the gem-cyclo three, well-separated sets of signals are expected, while for the nongeminal substitution, two isomers can be obtained: the 2,4,6-cis with one singlet signals and the 2,4,6-trans two singlets (or in some cases multiplets) corresponding to a AX_2 system. ^{31}P NMR data for similar structures¹⁹⁻²⁵ are also in agreement with that of **(2)**.

This confirms the structure depicted in Chart 1 for compound **(2)**. In the ^1H -NMR spectrum the expected signals of the aminopropyl(triethoxy)silane group were observed (see experimental part). The signals of the groups $\text{NCH}_3(\text{CH}_2)_3\text{CN}$ (not showed in the experimental part) are normal and appears in the expected range¹⁷. In their IR spectrum the respective bands corresponding to the $\nu(\text{PN})$ ring, $\nu(\text{Si-O})$ and $\nu(\text{NH})$ vibrations were observed at frequency values similar to that of

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3 compound **(1)**. The $\nu(\text{CN})$ band of the group $\text{N}(\text{CH}_3)(\text{CH}_2)_3\text{CN}$ was observed at 2248 cm^{-1} .
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5 Additionally the mass spectrum of **(2)** did not shows m/e peaks corresponding to another substitution
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7 product other than that of **(2)**. The molecular ion was observed as expected (see experimental part) as
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9 well as another peaks arising from the loss of $\text{Si}(\text{OCH}_2\text{CH}_3)_3$, OCH_2CH_3 and $\text{N}(\text{CH}_3)(\text{CH}_2)_3\text{CN}$
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11 groups.

12 The reaction of $\text{N}_3\text{P}_3\text{Cl}_6$ with 3 equivalent of $\text{H}_2\text{N}(\text{CH}_2)_3\text{Si}(\text{OEt})_3$ in presence of triethylamine in
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14 toluene as a solvent yields, at reflux, the intermediate $\text{N}_3\text{P}_3[\text{NH}(\text{CH}_2)_3\text{Si}(\text{OEt})_3]_3[\text{Cl}]_3$ which reacts
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16 with $\text{HOC}_6\text{H}_4(\text{CH}_2)\text{CN}$ in acetone and in presence of K_2CO_3 to give $\text{N}_3\text{P}_3[\text{NH}(\text{CH}_2)_3$
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18 $\text{Si}(\text{OEt})_3]_3[\text{OC}_6\text{H}_4(\text{CH}_2)\text{CN}]_3$ **(3)** as a red-brown oil. Their ^{31}P -NMR spectra exhibit, similar to those
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20 of **(2)**, two signal at 20.89 (quartet) and 18.69 (triplet) ppm, with relative intensity of 2:1 typical of an
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22 AB_2 system see supporting information S2. ^1H -NMR spectrum was normal and similar (with
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24 respect to the aminopropyl triethoxy silane group signals) to that of **(2)**.

25 As previously discussed, this confirms nongeminal substitution and supports the structure
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27 proposed in Chart 1. In their respective IR spectra, the bands corresponding to the $\nu(\text{PN})$ ring, $\nu(\text{Si}-$
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29 $\text{O})$ and $\nu(\text{NH})$ vibrations were observed at frequency values similar to that of compound **(1)** and **(2)**.
30
31 Similarly to **(2)**, the $\nu(\text{CN})$ band of the $\text{OC}_6\text{H}_4(\text{CH}_2)\text{CN}$ group was observed at 2240 cm^{-1} . The mass
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33 spectrum showed the expected molecular ion as well as another peaks arising from the loss of the
34
35 fragments OCH_2CH_3 , $\text{Si}(\text{OCH}_2\text{CH}_3)_3$ and $\text{OC}_6\text{H}_4\text{CH}_2\text{CN}$.

36 37 38 **Pyrolysis of the Si precursors**

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41 Pyrolysis of the silicon-containing cyclic trimer phosphazenes **(1)**, **(2)** and **(3)** results in gray
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43 solids in yields of 20-30%. Morphologies of the products depend on the temperature of the pyrolysis
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45 and on substituents around the phosphazenes phosphorus. For instance, in comparing the morphology
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47 (SEM) of the pyrolytic products at 800°C from **(1)**, **(2)** and **(3)**, shown in Fig. 1, we observe
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49 spherical shapes from **(1)**, mostly irregular shapes from pyrolysis of **(2)** and porous materials from
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51 **(3)**. EDAX analysis exhibits the presence of silicon, phosphorus and some oxygen atoms. A
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53 representative EDAX spectrum is shown in Fig. 1d.
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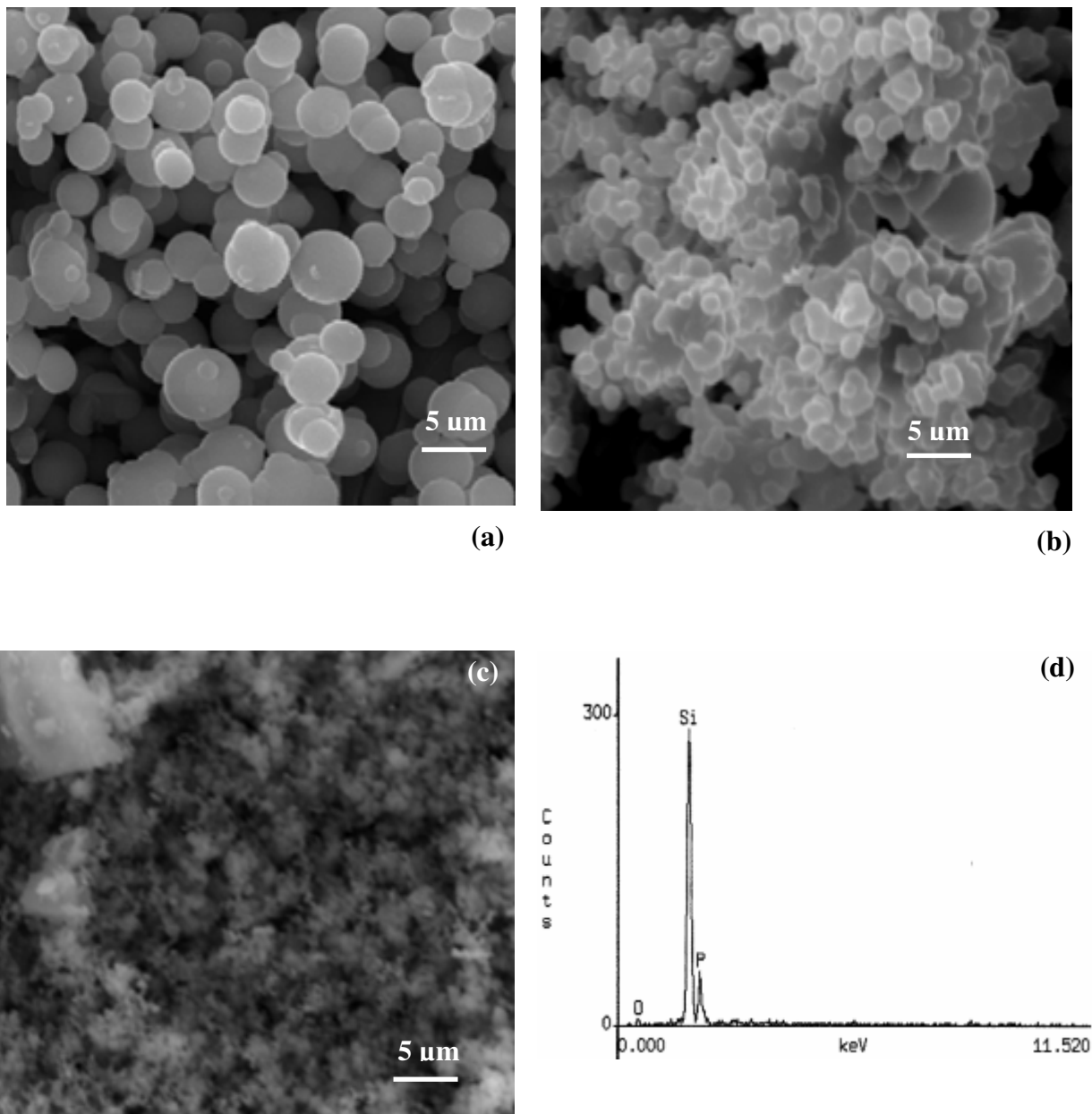
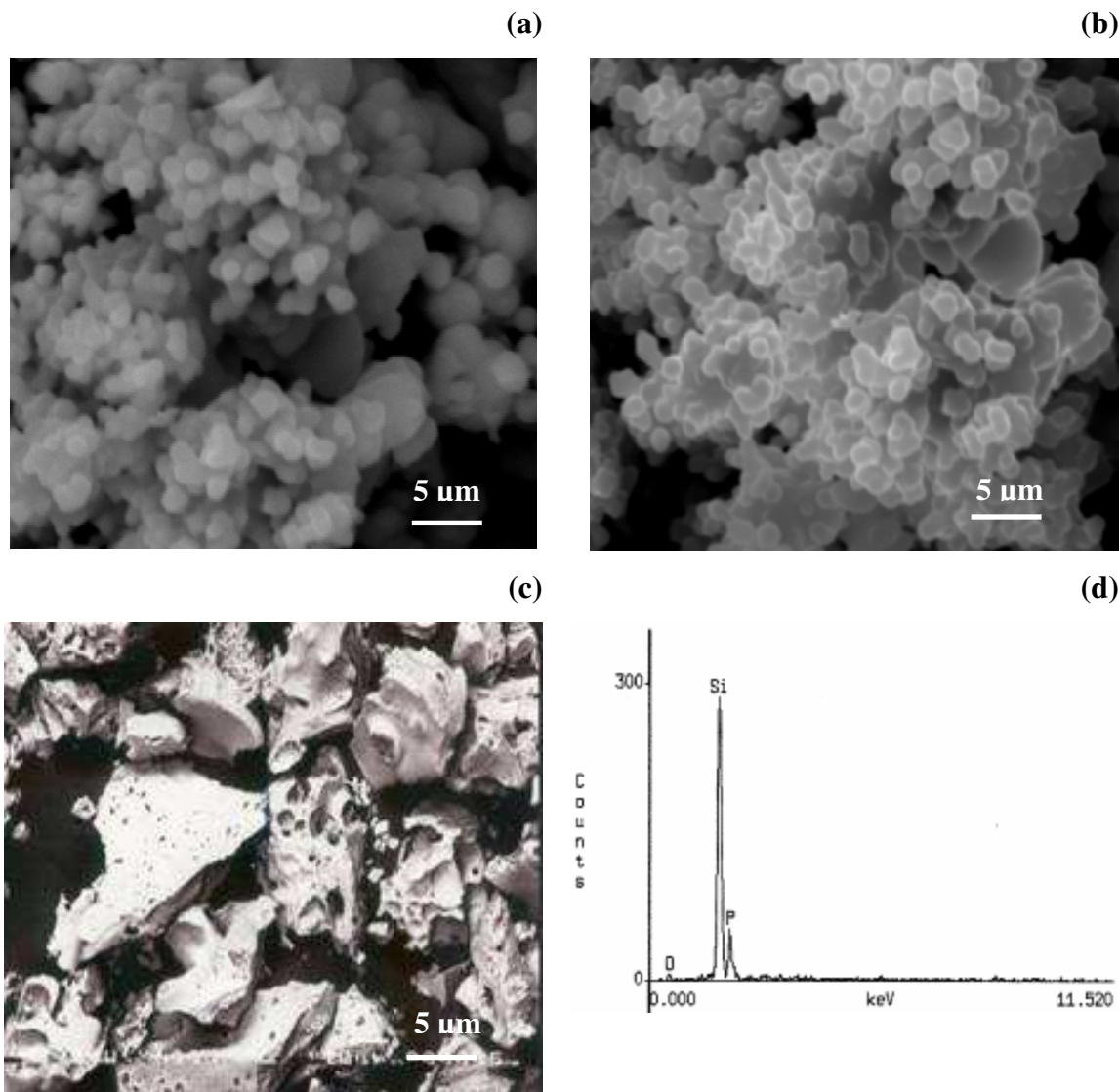


Figure 1 SEM images of the pyrolytic products from precursors (a) (1), (b) (2) and (c) (3). A representative EDAX spectrum is also shown in (d).

The temperature can also affect the morphology as is seen in Fig. 2 for the products from pyrolysis of (2) at 600°C, 800°C and 1000°C. At the highest temperature, a dense and ceramic-like structure is observed. Composition of the materials was investigated by EDAX analysis, powder X-ray diffraction and IR spectroscopy. EDAX analysis for the pyrolysis products from (1), (2) and (3)

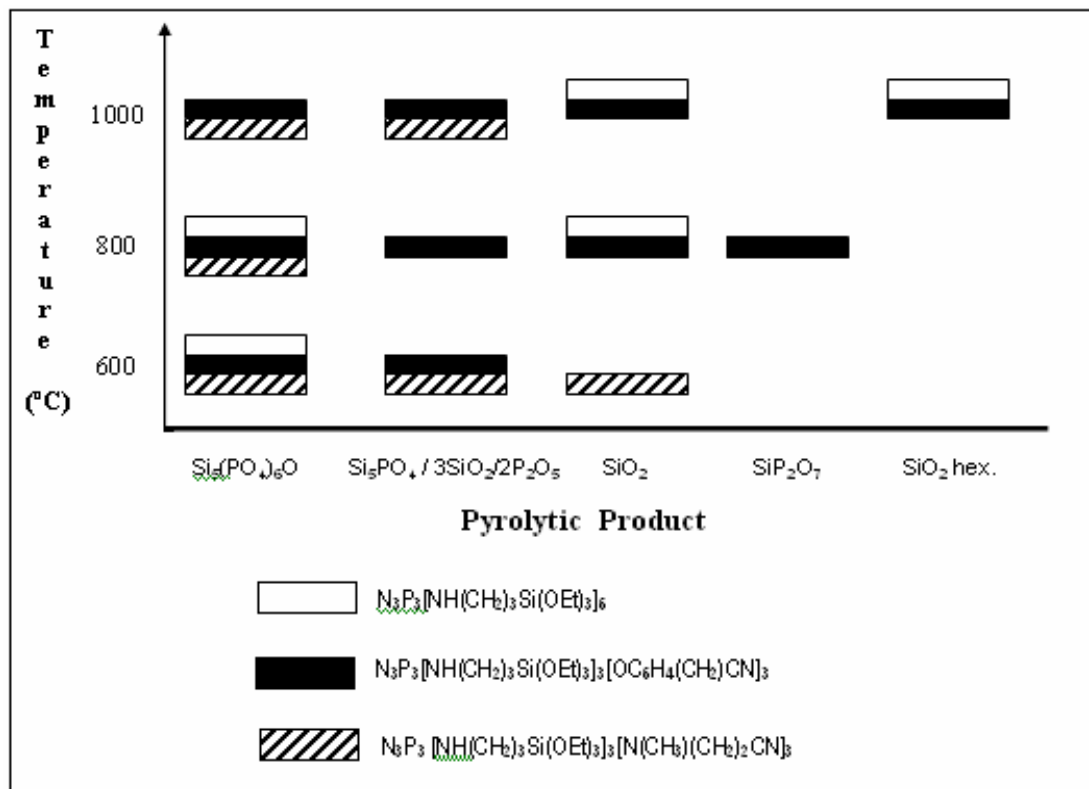
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3 exhibits similar patterns confirming the presence of silicon, phosphorus and trace quantity of oxygen
4 as is shown in Figs 1d and 2d.
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45 **Figure 2** SEM images of the pyrolytic products from precursor (**2**) at several temperatures: (a) 600
46 °C, (b) 800°C and (c) 1000°C. A representative EDAX spectrum for these products is also shown in
47 (d).
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52 The powder diffraction patterns of the products, shown in Fig. 3 are somewhat dependent on the
53 temperature as well as on the nature of the OR groups linked to the cyclotriphosphazenes core.
54 Scheme 2 shows the distribution of products after pyrolysis as a function of temperature for
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precursors (1), (2) and (3). For example, Fig. 3 shows the XRD pattern for (2) at 1000°C, (3) at 800°C and (1) at 1000°C. The small angle diffraction pattern for (1) at 1000°C is also shown.



Scheme 3 Schematic representation of the distribution of products from pyrolysis of precursors (1), (2) and (3) at several temperatures.

In general, three products from the reaction of SiO_2 and P_2O_5 were observed i.e., crystalline $\text{Si}_5(\text{PO}_4)_6\text{O}$ and SiP_2O_7 and a glassy amorphous phase $\text{Si}_5(\text{PO}_4)_6\text{O}/3\text{SiO}_2 \cdot 2\text{P}_2\text{O}_5$. It is known that the reaction of SiO_2 and P_2O_5 or SiO_2 with H_3PO_4 results in the formation of $\text{Si}_5(\text{PO}_4)_6\text{O}$, SiP_2O_7 or the glassy $\text{Si}_5(\text{PO}_4)_6\text{O}/3\text{SiO}_2 \cdot 2\text{P}_2\text{O}_5$ depending on the temperature and on the conditions of the reaction^{26a-e}. However the as obtained materials are microcrystalline sample and not nanostructured as the obtained by the here presented results.

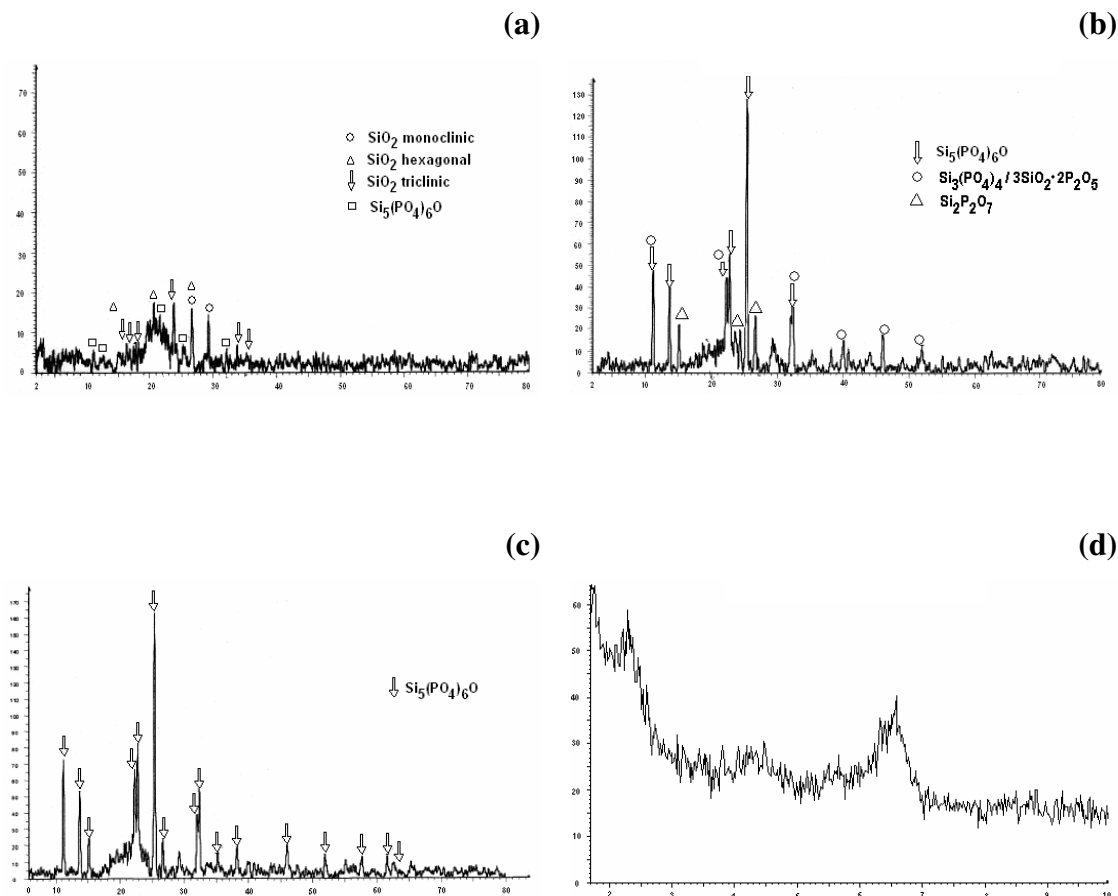


Figure 3 X-ray powder diffraction data for the pyrolytic products from: (a) **(1)** at 1000°C, (b) **(3)** at 800°C and (c) **(2)** at 800°C. (d) SAXRD pattern for **(1)** at 1000. °C.

The X-ray diffraction pattern for **(3)** at 800°C exhibits peaks corresponding to $\text{Si}_5(\text{PO}_4)_6\text{O}^{26b}$ (arrows), the glass $\text{Si}_5(\text{PO}_4)_6\text{O}/3\text{SiO}_2\cdot 2\text{P}_2\text{O}_5$ ²⁶⁻²⁸ (circles) and SiP_2O_7 ^{26c,d} (triangles), see Fig. 3b In contrast, the diffraction data for **(2)** at 800°C exhibits only peaks corresponding to $\text{Si}_5(\text{PO}_4)_6\text{O}$, Fig. 3c. Pyrolysis of **(1)** at 1000°C, however, results in several crystalline forms of SiO_2 as is shown in Fig. 3a. This is in good agreement with the previously reported decomposition of SiP_2O_7 at temperature above 1000°C^{26b-d}, to give SiO_2 and P_2O_5 . Some hexagonal MCM-41 SiO_2 phase was

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3 also formed as is seen from the low angle XRD pattern shown in Fig. 3d, similar to a recent report of
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5 SiP_2O_7 grafted to MCM-41^{26f}.
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7 From Scheme 1, it can be seen that $\text{Si}_5(\text{PO}_4)_6\text{O}$ is present in the pyrolysis of precursors **(1)**, **(2)**
8 and **(3)** at all temperatures assayed. In contrast, the pyrophosphate salt SiP_2O_7 was present only for
9 precursor **(3)** at 800°C. The glass of composition, $\text{Si}_3\text{PO}_4/3\text{SiO}_2/2\text{P}_2\text{O}_5$ was also observed in the
10 pyrolysis of **(3)** at the all temperatures examined. This product was only observed at 600°C and
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12 1000°C for the pyrolysis of **(3)**. Conversely, SiO_2 was obtained for all three precursor but at only at a
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14 single temperature for each one. Amorphous and crystalline materials of composition $\text{M}_x\text{O}_y/n\text{SiO}_2/$
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16 $m\text{P}_2\text{O}_5$ are interesting due to third-order nonlinear optical susceptibility properties, when M_xO_y is a
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18 transition metal oxide²⁷, and bioactive ceramic behavior is observed when the composition M_xO_y is
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20 that of Na_2O , MgO or CaF_2 ^{28,29}.
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23 The microspheres formed in the pyrolysis of precursor **(1)** at 800°C are the first generated
24 using phosphazenes as a template¹⁰. The resulting spheres were further investigated by means of
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26 TEM. Figure 4a-f displays the micrographs of a sample after calcinations at 800°C.
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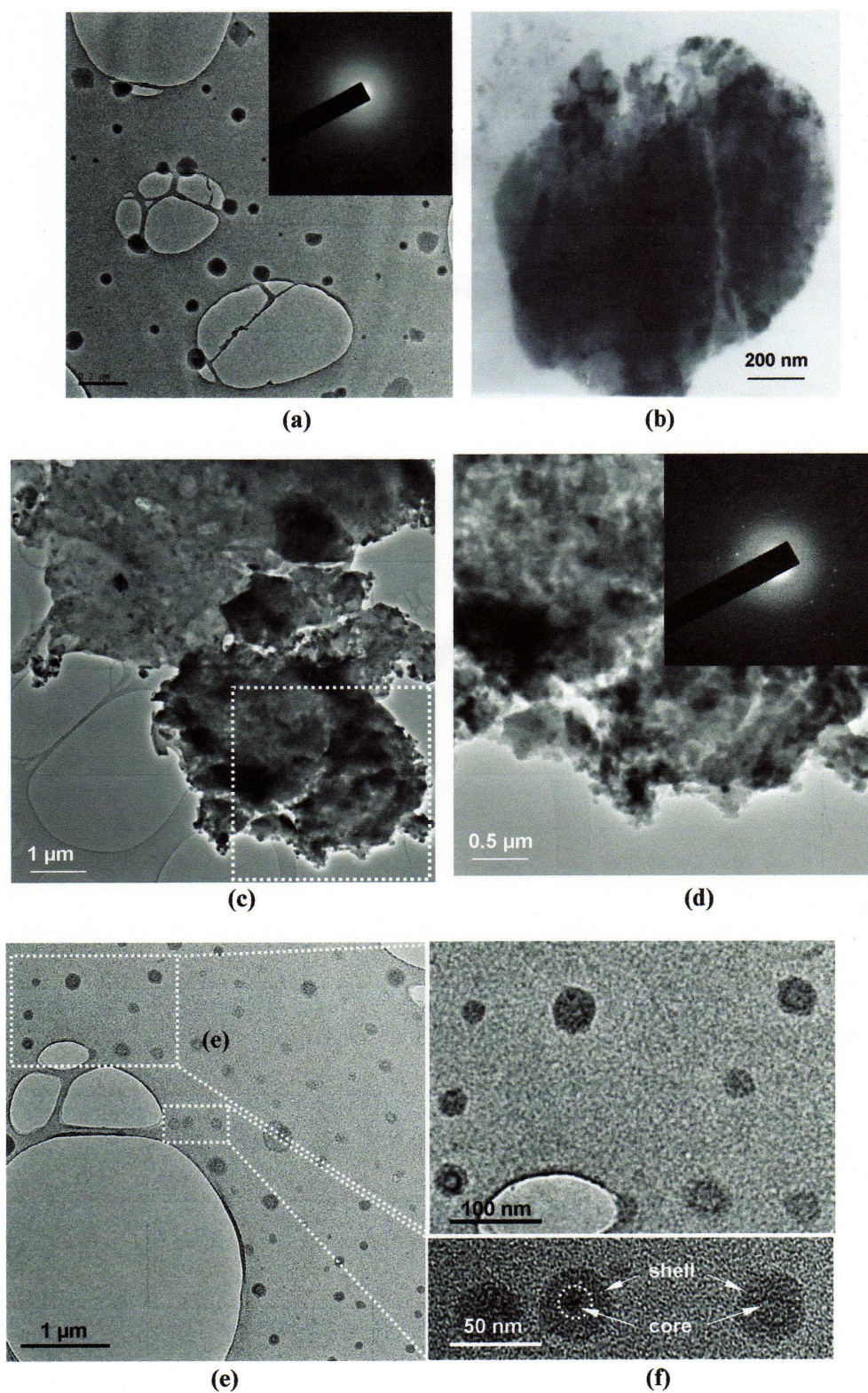


Figure 4 TEM images of pyrolytic products from (a), (b) (**1**) ; (b), (c) (**2**), and (e), (f) (**3**) at 800°C. Inset to (b) and (d) are the respective electron diffraction images patterns .

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3 The core-shell spheres structure of the aggregate nanoparticles of $\text{Si}_5(\text{PO}_4)_6\text{O}/\text{SiO}_2$ was possible to
4 observe clearly in the case of pyrolysis of precursor **(3)**, as is shown in figures 4e,f. For all the cases
5 the core-shell structure was also corroborated from energy dispersive X-Ray spectroscopy data and
6 using a method previously reported^{30,31} based on the core/shell contents in this case, from the Si/P
7 contents at of SiO_2 spheres coated with $\text{Si}_5(\text{PO}_4)_6\text{O}$. As is illustrated for the pyrolytic product from
8 **(1)** and **(2)** the electron diffraction image indicates somewhat amorphous material in all the cases.
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17 **BET Studies**

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21 BET studies were conducted on compounds **(1)**, **(2)** and **(3)** to quantify the degree of porosity
22 observed by SEM. The N_2 adsorption isotherms of the pyrolytic product from **(1)** indicate the
23 presence of mesopores with a mean diameter of 24.4 nm as measured from the pore size distribution
24 curve with a relatively low surface area of $3.2 \text{ m}^2 \text{ g}^{-1}$. Similar observations are found for the pyrolytic
25 products from **(2)** and **(3)** with values of mean pore size of 52.1 nm and 19.9 nm, respectively, with
26 corresponding surface areas of 2.0 and $2.3 \text{ m}^2 \text{ g}^{-1}$. These values are less than those found for
27 $\text{SiO}_2 \cdot \text{P}_2\text{O}_5 \cdot \text{M}_x\text{O}_y$ glasses³³.
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35 **Possible formation mechanism of the Si nanostructured materials**

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38 Some insight into the mechanism of formation of the nanostructured silicon materials from
39 the silylated cyclotriphosphazenes precursors can be obtained from the TGA/DSC in air, shown in
40 Fig. 5a. TGA and DSC for the pyrolytic products from **(2)** and **(3)** exhibit similar behaviour see
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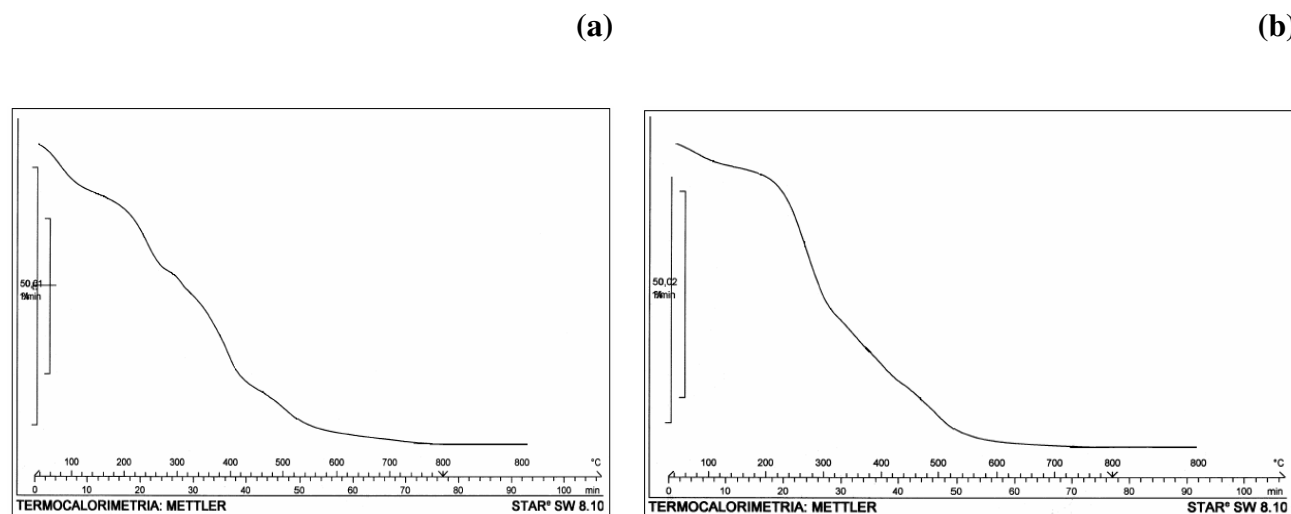
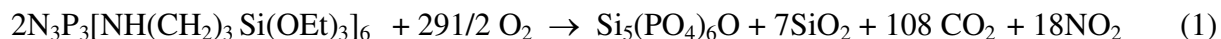


Figure 5. TGA curve for the pyrolytic product from precursor **(1)**: (a) in air and (b) in N_2 supporting information S5-S11

The initial weight loss of 9.31% is attributed to the volatilization of NO_2 by oxidation of the nitrogen content of the cyclic trimer. The second weight loss corresponds to the carbonization of the organic matter arising from the $N(CH_3)(CH_2)_3CN$ groups. The following weight loss is attributed to loss of CO_2 from carbonization of the remainder organic matter, i.e. the $Si(OEt)_3$ groups. The 41.19% pyrolytic residue is close to that expected for SiO_2 , $Si_5(PO_4)_6O$ 39.76%. It is at this point that the cyclotriphosphazene acts as a hybrid organic-inorganic template in the solid state in the formation of the silicon nanoparticles. This implies that on initial heating, a cross-linked structure involving cyclotriphosphazene linked by Si-O-Si bridges is formed. In fact organosilicon derivatives of cyclotriphosphazene undergo cross-linking of the ring on heating³⁴. The organic moiety, after calcination, produces holes in the cross-linked structures, which permit agglomeration of the silicon containing particles. The inorganic PN backbone of the polyphosphazenes in the presence of oxygen provides phosphorus atoms for the formation of the corresponding phosphorus oxides which react with the SiO_2 to form $Si_5(PO_4)_6O$ (or $Si_5(PO_4)_6O/3SiO_2 \cdot 2P_2O_5$ or SiP_2O_7 in some cases).

DSC curves also performed in air (see Supporting Information S9-S11) are also in agreement with this proposed mechanism. In fact, the exothermal peaks at 250°C and 300°C can be assigned to the carbonization of the organic matter. Similar exothermic peaks have been observed during the oxidation of organometallic complexes³⁶⁻³⁸.

Interestingly, the TGA curve in N_2 (see Fig. 5b) exhibits similar loss weight patterns. According to the equation for the combustion of the trimer **(1)** :



the oxygen content necessary for the formation of SiO_2 and $\text{Si}_5(\text{PO}_4)_6\text{O}$ in the absence of air (N_2 in this instance) can be partially supplied by the oxygen from the $\text{Si}(\text{OEt})_3$ groups, but the quantity available is insufficient for the total carbonization of the organic matter.

Thus, this mechanism is in agreement with those proposed for the formation of nanostructured metal foams from thermal decomposition of bi(tetrazolato)amine complexes³⁸ and from thermal decomposition of $[\text{Fe}(\eta\text{-C}_5\text{H}_4)_2(\text{SiRR}')_n]$ polymers³⁹.

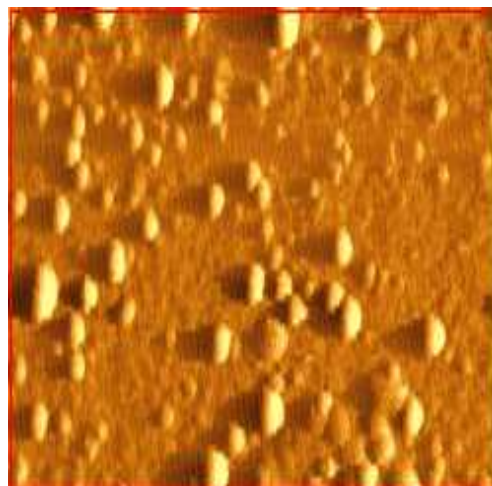
IR spectra

The IR and Raman spectra of the products from the reactions of SiO_2 with P_2O_5 have been studied in detail⁴⁰. All pyrolytic products from **(1)**, **(2)** and **(3)** exhibit absorptions at 1180-1050 (vs, broad), 800-680 (w), and 490 (s, broad), which are typical of some of the reaction products of SiO_2 with P_2O_5 , i.e. crystalline $\text{Si}_5(\text{PO}_4)_6\text{O}$, SiP_2O_7 or the amorphous glass $\text{Si}_5(\text{PO}_4)_6\text{O}/3\text{SiO}_2 \cdot 2\text{P}_2\text{O}_5$.

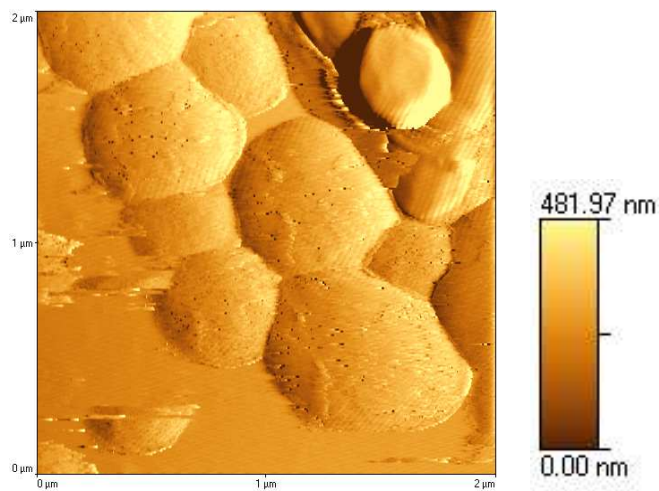
Surface morphological studies of deposited precursors

Suitable AFM studies of the trimer samples were achieved by dissolving a dichloromethane solution followed by dropwise deposition onto a silicon wafer, evaporation at room temperature and pyrolysis at 800°C. For trimer **(1)** separated grains of average height ~20 nm were observed as is shown in Fig. 6.

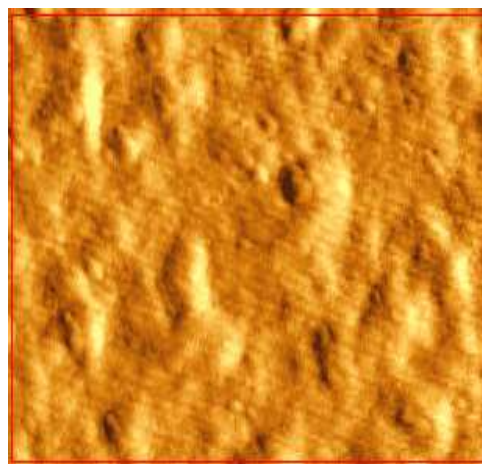
The pyrolytic product from trimer **(2)** is observed to form large islands with varied size (from 50 nm to 180 nm) that nucleate and grow progressively in three dimensions, as is shown in Fig. 6b. In several cases these islands are either isolated or joined forming a chain, shown in Fig. 6b. For the pyrolytic product from trimer **(3)**, a more uniform coverage of the nanostructures is observed, as is shown in Fig. 6c.



(a)



(b)



(c)

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3 **Figure 6.** AFM images of the pyrolytic products from (a) precursor **(1)**, (b) precursor **(2)** and (c)
4 precursor **(3)**.
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8 **Conclusions**

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11 Siliceous nanostructured materials are accessible in a wide range of morphologies and
12 composition through the solid-state pyrolysis of new precursors **(1)**, **(2)** and **(3)**. Both the morphology
13 and composition of the nanostructured products are strongly dependent on the nature of the groups
14 around the phosphazenic ring and on the temperature of the pyrolysis. Oxidation of the phosphorus of
15 the polymeric chain and of the silicon from the siloxane groups gives P_2O_5 and SiO_2 , respectively,
16 which react resulting in silicon phosphates in their crystalline phases $Si_5(PO_4)_6O$ and SiP_2O_7 and/or
17 amorphous phase as the glass $Si_5(PO_4)_6O/3SiO_2 \cdot 2P_2O_5$. These nanoparticles growth inside the
18 vacancies, which are formed by carbonization of the organic matter, given rise to nanostructured
19 silicon compounds.
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28 AFM images evidence increasing grains density on the surface of the deposition on going to pyrolytic
29 precursor **(1)** to **(3)**. This new technique shows promise for being a flexible, general approach to the
30 formation of a range of nanostructured siliceous pyrophosphate materials not currently accessible by
31 other methods.
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40 Universidad de Chile) for the BET measurements and to Dr. Calum Dickinson for the TEM
41 measurements. Financial support from FONDECYT, project 1085011, is also acknowledged.
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46 **Supporting information available:** ^{31}P NMR spectrum for **(1)**, **(2)** and **(3)** and TGA curves for
47 **(2)** and **(3)** in air and N_2 and DSC curves for **(1)**, **(2)** and **(3)** in air and nitrogen.
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