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Fabrication of transparent composites for non-invasive Surface Enhanced Raman Scattering (SERS) analysis of modern art works

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Abstract

Surface Enhanced Raman Scattering (SERS) has become an invaluable tool for the identification of colorants in artworks, due to its enhanced sensitivity and ability to quench fluorescence interference compared to Raman spectroscopy. However, the application of SERS to artwork analysis is still limited by its inherent invasiveness and the need for extraction procedures. In this work non-invasive transparent SERS probes were fabricated by self-assembly of Ag nanoparticles into glass and PDMS surfaces and used for identification of dye content in artistic drawings. SERS measurements were performed directly on the selected analytical surfaces by laser back illumination through the SERS probe. The non-invasiveness of fabricated probes was tested by optical microscopy. It was found that Ag nanoparticle/glass probes left no Ag nanoparticle residue after four consecutive depositions on sacrificial surfaces, whereas Ag nanoparticle/PDMS composites could be deposited and subsequently peeled off the analytical surfaces leaving no contamination traces. The high conformability, flexibility and transparency of Ag nanoparticle/PDMS composites enabled good adhesion to the surface of analyzed artistic drawings and therefore the generation of in situ SERS spectra from artistic drawings. The use of this method allowed identification of main dye components in real artworks comprising a red BIC ballpoint drawing and a Japanese woodblock print.

Keywords: Raman spectroscopy, SERS, Non-invasive, Paper artwork, BIC ballpoint pens, Japanese wood print

Introduction

The ability to identify colorants and colorant mixtures in artworks is of fundamental importance to inform application of conservation treatments, display and preservation procedures as well as to allow identification of artworks' provenance and forgery. Due to its inherent non-invasive nature, fingerprinting capability, easiness of use and availability of portable instrumentation, Raman spectroscopy is among the most versatile and used spectroscopic techniques for analysis of artworks [1–3]. However, the low sensitivity and the strong fluorescence interference of colorants in the visible part of the spectrum, have largely limited the application of Raman to the analysis of natural and artificial colorants. In the last 20 years Surface

Enhanced Raman Scattering (SERS) has been established as powerful tool for the identification and analysis of natural and synthetic colorants used in artworks [4–6]. The high sensitivity, up to 10^8 – 10^{12} times the sensitivity of Raman spectroscopy and down to single molecule detection, combined with the retained fingerprinting capabilities of Raman spectroscopy, offers key advantages for the identification of colorants, often present in artworks in small concentrations, mixed with other colorants and embedded in complex matrices [7]. In addition, SERS can effectively quench the fluorescence interference caused by laser illumination, thus overcoming Raman spectroscopy limitations and considerably widening the range of applications in conservation science. Accordingly, SERS has been successfully used for investigations of a large range of heritage materials including archeological textile fibers [6], woodblock prints [8, 9] and has enabled identification of synthetic and organic dye content in

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polychromic works of art [10, 11], painting samples [12, 13] oil paintings [14] and commercial inks [15]. Another advantage of SERS is the availability of lightweight, sensitive and low cost portable instrumentation, whose use is often mandatory for the analysis of large or unmovable artworks.

SERS experiments are usually performed with the use of colloidal concentrated Ag solutions or with rigid ordered plasmonic substrates. Both methods are inherently invasive as they cannot be used directly on the analytical surface and require the use of harsh extraction methods in order to separate by hydrolysis the colorant from its matrix, before application of Ag colloids or deposition onto Ag rigid SERS substrates [16]. Alternative use of Ag colloidal pastes has been proposed for in situ, non-hydrolysis SERS measurements [17–19]. This approach consists into deposition of a droplet of concentrated colloidal solution on untreated samples and has been shown particularly effective for the characterization of historical textiles. However, the method is still inherently invasive as direct deposition of Ag pastes stains the work of art, making necessary the physical removal of analyte fragments and their immobilization onto arbitrary substrates prior Ag paste application.

Consequently, many efforts in the last few years have concentrated on the development of non-invasive SERS substrates. A micro-extraction process was developed by Leona et al. who used a bead of polymer hydrogel loaded with a water/DMF/EDTA mixture to extract minimal amounts of colorants from works of arts [20]. The hydrogel with extracted colorants was then treated with Ag nanoparticles for SERS analysis. Successful demonstration of non-invasive identification of crystal violet on deep purple colored areas of Japanese woodblock prints on paper was obtained by this method. Doherty et al. designed and fabricated an active SERS film constituted by Ag nanoparticles doped into a methylcellulose matrix [21]. The resulting composite was transparent, could be applied to the surface of an artwork and removed after analysis. Strongly enhanced SERS spectra were obtained for powder pigments and lake/egg tempera, otherwise not measurable by Raman spectroscopy. However, close observation of detached composites showed that analytical surface fragments (40 μm) underwent removal and trapping into the composite matrix, making the method slightly invasive in nature. One step towards truly non-destructive SERS analysis is the implementation of “dry state” measurements, which has been achieved by immobilization of Ag colloids on the distal end of an optical fiber. In this way Polwart et al. successfully obtained SERS spectra of crystal violet from historical costumes of the late nineteenth century [22]. Stokes et al. fabricated a SERS-inducing nanoprobe by thermally evaporating a

thin layer of silver nanoparticles to the tip of a tapered optical fiber (100 nm tip diameter) [23]. Spectra of brilliant cresyl blue coated on a glass slide were recorded by touching dry dye samples with the nanoprobe. Simpler substrates were fabricated by Zaffino et al. by deposition of Ag islands (fabricated by beam evaporation) or Ag colloidal particles (chemically synthesized) over glass slides [24]. Spectra of alizarin, purpurin and lac dye were obtained by deposition of Ag island substrates on dyed wool threads. However, the authors report that scanning electron microscopy energy dispersive X-ray analysis (SEM-EDX) of post-treated wool threads showed contamination from Ag islands and colloidal particles. Although residues were small in size (1–3 μm), the oxidation of Ag may lead to surface color changes not acceptable for works of art analysis.

The aim of this work was to overcome the limitations above through the fabrication of novel SERS substrates obtained by deposition of self-assembled Ag nanoparticles into glass and PDMS substrates and to demonstrate their application in non-invasive analysis of artworks.

Experimental section

Materials and methods

Silver nitrate, sodium borohydride, trisodium citrate, octadecylamine (ODA), chlorobenzene, methanol, ethanol were purchased from Sigma-Aldrich used without further purification. Reference dyes Crystal violet (CV), Rodhamine B and Blue 38 were also purchased from Sigma-Aldrich. SYLGARD 184 Silicon Elastomer Kit for polydimethylsiloxane (PDMS) film fabrication was also purchased from Sigma-Aldrich. All glassware was cleaned with *aqua regia* prior to silver nanoparticles synthesis. Milli-Q water (resistivity 418 $\text{M}\Omega/\text{cm}$) was used throughout the experiments.

Medium and Crystal BIC pens were purchased from local stores. Japanese prints and Anne-Flore Cabanis drawings were kindly donated by paper conservator Antonio Mirabile.

Synthetic procedures

Ag nanoparticle synthesis

Ag nanoparticles were synthesized according to a process reported previously by Rainville et al. [25]. Briefly, 12 mL of an aqueous solution (0.2% w:w) of silver nitrate and 488 mL of deionized water were mixed and heated to 100 °C. Next, 11.6 mL of an aqueous solution of sodium citrate (3.4×10^{-2} M) was added, followed 30 s later by the quick injection of 5.5 mL of a freshly prepared, ice-cold aqueous solution of sodium borohydride (0.2 M) and sodium citrate (0.03 M, 50 mL). The solution was stirred for 2 min and subsequently left it to cool at room temperature. The obtained nanoparticle solution was

centrifuged once at 9000 rpm for 20 min and redispersed in 4 mL of distilled water. The concentrated nanoparticle solution was stored in the dark for subsequent use.

Ag nanoparticles were transferred to organic phase by slight modification of a method previously reported [26]. Briefly, 2 mL of the concentrated aqueous Ag nanoparticle solution were added to 2 mL of a 2×10^{-5} M solution of ODA in chlorobenzene. The two solutions were vigorously stirred for 60 min to facilitate the transfer of Ag nanoparticles to the organic phase.

Ag nanoparticles/glass/PDMS nanocomposites fabrication

PDMS was fabricated by pouring 10:1 (w:w) of Sylgard 184 elastomer and curing agent in a petri dish followed by heating at 60 °C for 24 h. After removal from oven, the PDMS film was rinsed with isopropanol several times, dried with nitrogen gun and cut to small squares of 1×1 cm. The final thickness of the PDMS film was 1 mm.

In order to obtain Ag nanoparticle/glass and PDMS composites, a small aliquot (10 μ L) of Ag nanoparticle chlorobenzene solution was deposited on an acetone-cleaned glass coverslip. The droplet was left in contact with air (5 min) until Ag nanoparticles spontaneously self-assembled at the solvent/air interface, as shown by the formation of a metallic lustre layer. A transparent film (either glass or PDMS) was used as support to retrieve the assembled nanoparticles by bringing it into contact with the nanoparticle droplet for 30 s. Excess organic matter was removed by immersing the Ag nanoparticle/glass/PDMS composite in isopropanol overnight, followed by multiple rinses with fresh isopropanol.

Characterisation

UV–vis spectra were acquired using an Agilent/HP 8453 UV–vis Spectrophotometer (200 nm to 1100 nm). Scanning electron microscopy (SEM) images of Ag nanoparticles and Ag nanoparticle/PDMS composite films were acquired using a field emission (JSM- 6700F, JEOL UK Ltd.) scanning microscope operating at 3 kV.

Raman and SERS measurements were obtained from a Renishaw inVia Raman system equipped with a 514 argon ion laser. The laser beam was focused onto the sample through a Leica 20X objective with 0.4 N.A. Acquisition time was usually 10 s and measured power was 3 mW. Hand held Raman spectra at 785 nm were obtained from an InPharma spectrometer. The laser power was 50 mW at sample and acquisition time was between 7 and 20 s.

SERS spectra of reference dyes were taken by back illumination of a slide containing a dried droplet of dye solution (10 μ L, 1 mM in MeOH) with Ag nanoparticle/glass composites. SERS measurements of colored papers and paper artworks were performed by gently pressing Ag nanoparticle/glass (18 $\times$ 18 mm) or Ag nanoparticles/

PDMS composites (10 $\times$ 10 mm) on the analytical surface (BIC colored paper or paper artwork) for few seconds. SERS signals were collected by back laser illumination. All spectra were background subtracted.

Results and discussion

Fabrication and characterization of SERS composites

Ag nanoparticles used in this study were prepared in aqueous solution by reduction of Ag(I) salts with strong reducing agent sodium borohydride and subsequently phase transferred in organic solution. SERS probes were constituted by dense layers of Ag nanoparticles immobilized on transparent rigid (glass) and flexible (PDMS) substrates from self-assembled chlorobenzene solutions. The use of organic solutions was motivated by previous studies carried out by our group and others showing that direct deposition of aqueous plasmonic colloidal solutions on glass substrates resulted in dis-homogeneous nanoparticle distribution (i.e. accumulation of nanoparticle along the perimeter of the droplet), due to coffee stain effects [27, 28]. In contrast, homogeneous arrays were obtained by surface transfer of self-assembled superstructures obtained from evaporation of organic droplets. The process of SERS probe fabrication is schematically shown in Fig. 1.

A small droplet (10 μ L) of concentrated Ag nanoparticle in chlorobenzene solution was deposited on a clean glass substrate and left undisturbed for 5–10 min (1a). Following the evaporation of small amount of solvent, Ag nanoparticles spontaneously self-assembled at the solvent/air interface, as evidenced by formation of a metallic lustre (1b). At this point a receiving substrate (glass or PDMS) was brought in contact with the nanoparticle droplet (1c). Capillary effects promoted the transfer of the metallic layer on the receiving substrate (1d). Following transfer, the glass/PDMS substrate containing Ag nanoparticles was allowed to dry and subsequently immersed overnight in isopropanol to remove unbound particles and excess surfactants present in the organic solution. Figure 1e, f show photographs of Ag nanoparticles deposited on glass and PDMS substrates, respectively.

Figure 2 show SEM images of (a) as-synthesized Ag nanoparticles in aqueous solutions showing formation of spherical particles with average size of 15 ± 5 nm and (b) Ag nanoparticles transferred on glass substrates from self-assembled chlorobenzene droplets, following the method described in Fig. 1. The images showed that particles retained their spherical size through the phase transfer/self-assembly/substrate transfer process and that the final average size of SERS substrate deposited particles was 20 ± 3 nm. More details on UV–vis response and

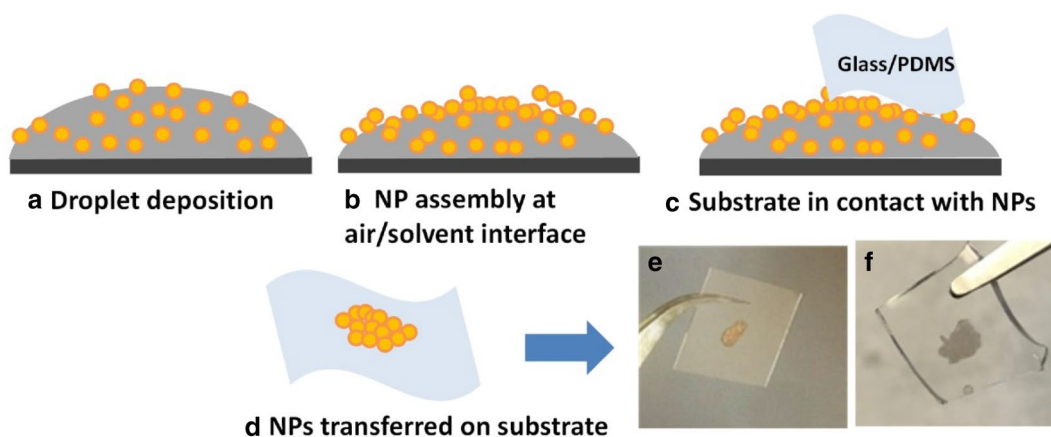


Fig. 1 **a–d** Schematic of SERS Ag nanoparticle/substrate (glass and PDMS) fabrication. Photographs of **e** Ag nanoparticle/glass and **f** Ag nanoparticles/PDMS composites

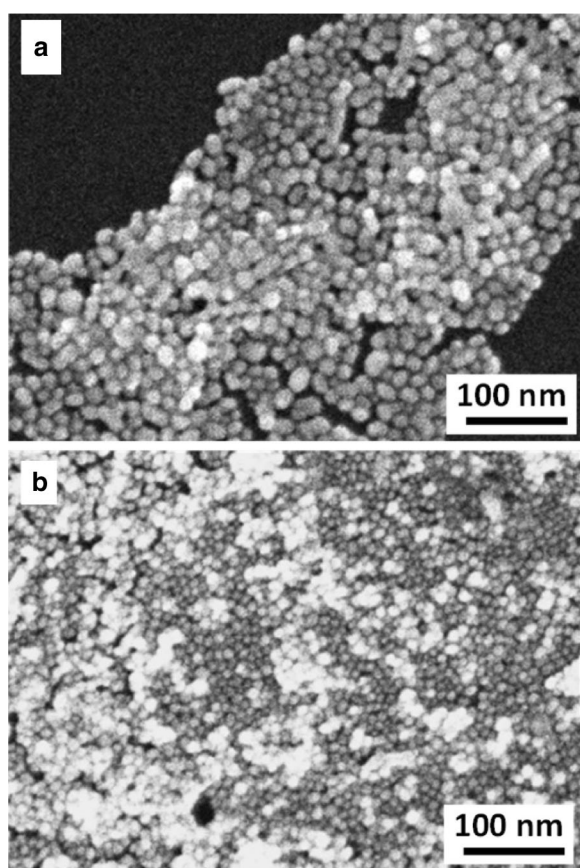


Fig. 2 SEM images of Ag nanoparticles transferred on **a** glass substrate and **b** PDMS substrate from self-assembled chlorobenzene droplets

size distribution are reported in Additional file 1: Figures S1, S2).

Minimal invasiveness tests

Prior attempting SERS analysis on real works of art tests were performed in order to establish the capability of fabricated SERS structures for minimal-invasive analysis. To this end, Ag nanoparticles/glass and Ag nanoparticle/PDMS substrates were repetitively placed for few minutes on marked areas of commercial white paper sheets. Close contact between the Ag nanoparticle/glass substrates and the analytical white paper sheet was obtained by placing two glass slides on the sides of the SERS substrate (see Fig. 4a). During SERS measurements this procedure ensured that the substrate was held in place and also maintained close contact between the Ag nanoparticle moiety and the analytical surface necessary for the SERS effect to occur. Optical microscopy images of white paper sheets were taken prior SERS probe deposition (Fig. 3a) and after each deposition (Fig. 3b–f) to assess the accidental release of loosely bound nanoparticles on the analytical surface. As shown by Fig. 3b, the first Ag nanoparticle/glass substrate placement (deposition 1) on the paper substrate left nanoparticle residues; Fig. 3 shows that four repetitive placements had to be performed before no residual traces of Ag nanoparticles could be detected on the paper sheet. In contrast, when Ag nanoparticle/PDMS substrates were placed on the paper no residual traces of Ag nanoparticles were observed already after the first placement (Fig. 3g, h). It is important to stress

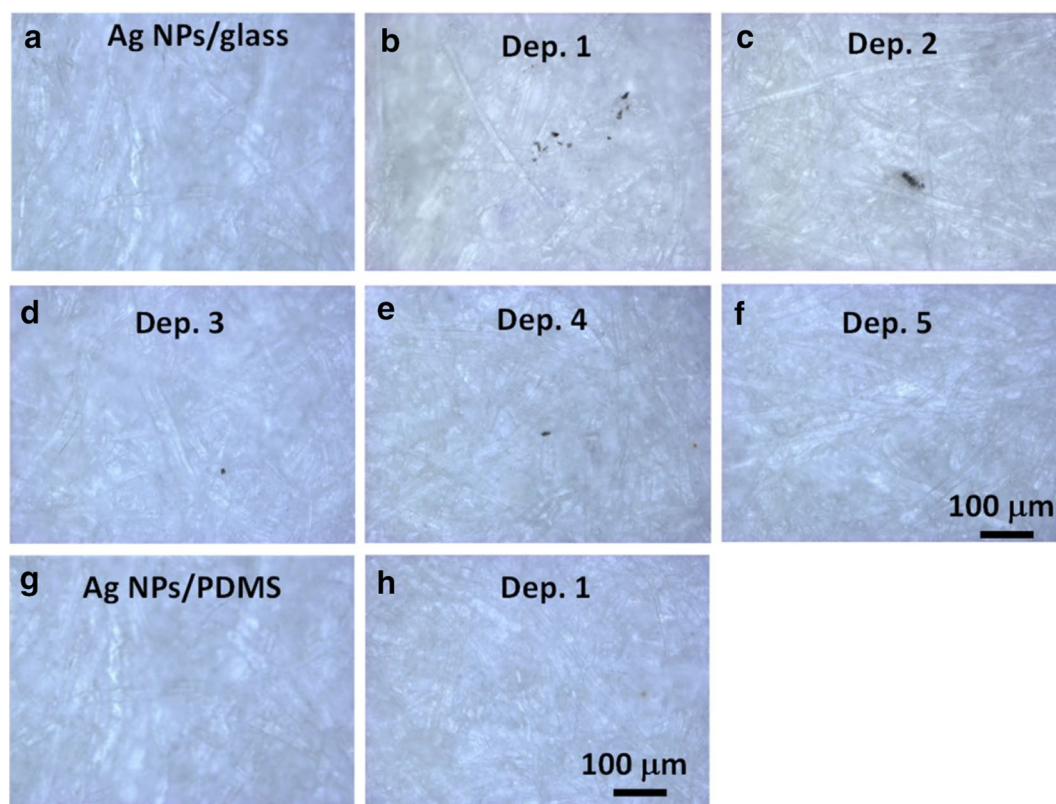


Fig. 3 **a** Photograph of Ag nanoparticle/glass SERS substrate; **b–f** optical microscopy images of a white paper sheet taken after consecutive depositions of the same Ag nanoparticle/glass SERS substrate showing presence of Ag nanoparticle residues for the first four depositions; **g** photograph of Ag nanoparticle/PDMS SERS substrate; **h** optical microscopy image of a white paper sheet taken after deposition of Ag nanoparticle/PDMS SERS substrate showing no trace residues of Ag nanoparticles

that Ag nanoparticle/PDMS substrates were applied to the analytical surface just by slight pressure. The substrates showed good adhesion with the paper surfaces and, in contrast with the Ag nanoparticle/glass substrates, did not require the use of additional weights to hold the SERS substrate in place or to maximize SERS effects. These data show that the adhesion of Ag nanoparticles to PDMS was stronger than the adhesion to glass substrates. One reason could be the strong hydrophobic affinity between PDMS and the hydrophobic ODA molecules crafted on the surface of Ag nanoparticles following their phase transfer in chlorobenzene. It should be pointed out that the strong adhesion of the Ag nanoparticle/PDMS composite to the paper analytical surface was enough to ensure close contact between the Ag nanoparticles and the analytical surface to promote the SERS effect. This effect was due to the use of PDMS and was promoted by the PDMS/

Ag nanoparticle area ratio in the composite. When transferred on PDMS the Ag nanoparticles covered an approximately circular area of 5 mm in diameter. However, the size of the used PDMS was larger (ca. four times in size). That means that the adhesion to the substrate was not provided by the nanoparticle area but by the PDMS area, which stuck to the surface and ensured adhesion of the overall composite.

SERS tests with Ag/PDMS composites

The capability of Ag nanoparticle/PDMS composites to perform SERS analysis was initially assessed by taking SERS spectra of model molecule 4-aminobenzene thiol (4-ABT). This particular molecule was selected the occurrence of a SERS effect is easily assessed by the occurrence of additional b_2 in plane, out of phase vibrational modes not present in the Raman spectrum and attributed to a SERS chemical enhancement effect (plasmonic metal–molecule charge transfer process) [29]. The spectrum of 4-ABT was taken by placing the composite on a dry droplet (1×10^{-6} M, MeOH) of 4-ABT on a glass

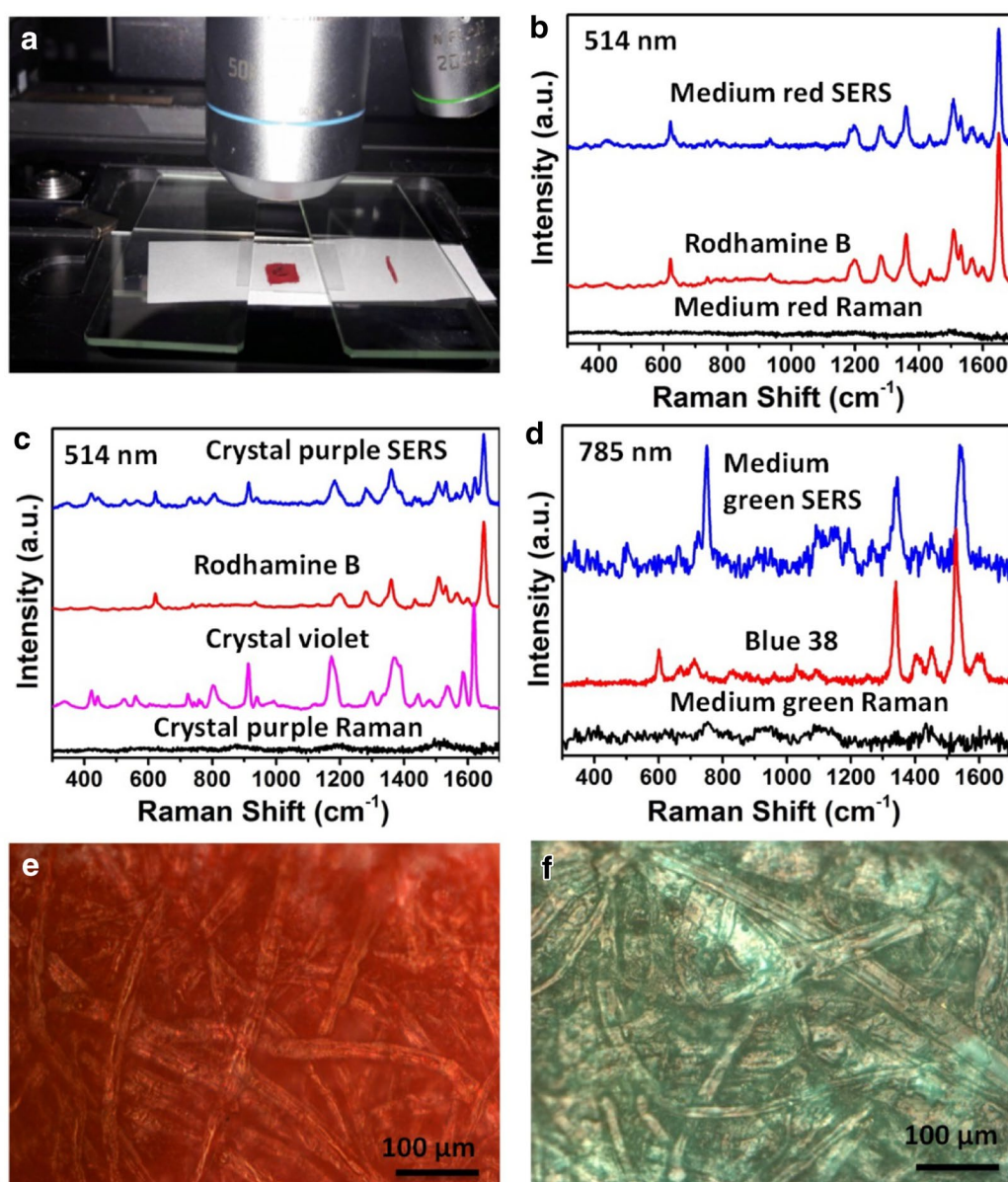


Fig. 4 **a** Photograph of set up used for SERS analysis of colored paper with Ag nanoparticle/glass SERS substrates; **b** SERS (blue line) and Raman (black line) spectra of BIC medium red pen square on paper, SERS spectrum of reference Rodhamine B dye (red line). Spectra were taken with bench Raman instrumentation at 514 nm illumination; **c** SERS (blue line) and Raman (black line) spectra of BIC crystal purple pen square on paper, SERS spectra of reference Rodhamine B (red line) and Crystal violet (pink line) dyes, 514 illumination; **d** SERS (blue line) and Raman (black line) spectra of BIC medium green pen square on paper, SERS spectrum of reference Blue38 dye (red line). Spectra taken with handheld Raman instrumentation at 785 nm illumination; **e–f** optical images of colored red and green colored squares after SERS analysis

slide with the Ag nanoparticle facing the 4-ABT. SERS spectra were obtained by back laser illumination and showed characteristics a_1 and b_2 modes of 4-ABT (see Additional file 1: Figure S5 for detailed description of the 4-ABT SERS spectrum obtained with Ag nanoparticle/PDMS composites).

SERS analysis on model ink-colored paper samples

In parallel, capabilities of fabricated substrates for in situ SERS analysis of commercial BIC ballpoint inks were tested. All spectra were taken in situ, by applying the Ag nanoparticle/glass substrates on pen colored squares with the Ag nanoparticle facing the analytical surface followed by direct back laser illumination (see Fig. 4a for

photographic details of the setup). All spectra were background subtracted. Figure 4b–d show Raman and SERS spectra of colored squares drawn on commercial paper by Medium red, Crystal purple and Medium green BIC ballpoint pens, respectively. A bench size Raman spectrograph with 514 nm excitation wavelength was used for analysis of the red and purple squares. Both colored squares showed featureless Raman spectra saturated by fluorescence. The generation of high fluorescence interference was not surprising as UV–vis spectra (see Additional file 1: Figure S3) showed that both red and purple colored inks were characterized by strong absorption centered at 554 nm, close to the selected excitation wavelength of 514 nm, which resulted in concomitant generation of interference fluorescence emission, and consequent masking of Raman signals. In contrast, the SERS spectra of red and purple colored squares showed dramatically enhanced spectral features (Fig. 4b, c blue lines). The red square showed clear presence of Rhodamine B (Pigment Violet 1, C.I. 45,170), evidenced from the presence of the following vibrational bands: 1647 (C–C bending and C=C stretching of xanthene aromatic ring), 1530, 1505 (aromatic C–H bending), 1398–1278 (aromatic C–C bending), and 620 (xanthene ring puckering) cm^{-1} [30]. Rhodamine B and/or rhodamine 6G had been previously identified in red ballpoint pen formulations through Easy Ambient Sonic-spray Ionization Mass Spectrometry (EASI-MS) analysis [31], and Laser Desorption/Ionization Time of Flight Mass Spectrometry (LDI-TOF-MS) [32]. Interestingly, while clear distinction between Rhodamine B and Rhodamine 6G could not been achieved with mass spectrometry techniques, the reported SERS analysis clearly showed presence of Rhodamine B in this particular Medium red BIC ink formulation (Fig. 4c, blue line). The SERS spectrum of the Crystal purple square displayed spectral features similar to the red square (main bands at 1650, 1505 and 1358 and 620 cm^{-1}), therefore also suggesting presence of Rhodamine B in the ink mixture. However, also small additional bands centered at 1623 (stretching of benzene rings), 938, 914 (asymmetric stretching and bending of C–C_{center}–C bonds), 565, 525, 443 and 420 (bending of the C–N–N bonds) cm^{-1} were observed, indicative of the presence of triarylmethane dye crystal violet (Methyl violet 10B, CI 42,555) in the ink mixture [33]. For comparison reference spectra of Rhodamine B and crystal violet are shown in Fig. 4b. The presence of crystal violet was further confirmed by the intense purple color of the ink, its occurrence in blue and black BIC pen ink formulations, and by previous Raman analysis performed on the Crystal purple ink at other excitation wavelengths [19]. It is important to stress that the blank SERS spectrum recorded by placing a Ag nanoparticle/glass substrate on a white

paper substrate was featureless and did not display any spectral feature attributable to the excess of stabilizing agents used to promote Ag nanoparticle phase transfer from aqueous to chlorobenzene phases (see Additional file 1: Figure S4).

Previous Raman analysis carried out on colored BIC pens showed that green inks possess good Raman response at 514 nm illumination wavelengths, making the application of SERS not necessary for identification and analysis [19]. However, in view of analytical restrictions imposed by the preciousness and limited mobility of works of art, often requiring the use of portable instrumentation, it is important to assess the capabilities of low-cost handheld Raman instrumentation for analysis of artistic drawings. Figure 4c shows the Raman and SERS spectra of a green colored square measured with a handheld Raman spectrograph with excitation at 785 nm. The Raman spectrum of the Medium green pen ink resulted featureless. However, the SERS spectrum showed strong enhanced spectral features with peaks centered at 1539, 1342, 758 and 499 cm^{-1} . From previous work carried out in our group, it is known that green inks are often constituted by a mixture of blue and yellow dyes and that blue phthalocyanine Blue 38 dye is used in BIC pen formulations [15]. In this case, identification of the yellow component could not be achieved with portable instrumentation. However, the presence of phthalocyanine Blue 38 was confirmed from peaks at 1344 and 1454 cm^{-1} , which were both attributed to internal vibrations of copper phthalocyanine (CuPh) macrocycle [19]. Figure 4e, f show optical images of colored papers taken after removal of SERS probes following analysis showing that no residue was left on the analytical surface.

SERS analysis on real works of art

Finally, Ag nanoparticle/glass and Ag nanoparticle/PDMS composites were used for SERS analysis of real works-of-art. Figure 5a show photograph of a drawing made in red BIC ballpoint pen by contemporary French artist Anne-Flore Cabanis. The peculiarity of the drawing was its overall circular pattern constituted by one continuous line bent at 90° angles. As the pen was never lifted from the paper during the drawing, the continuous line showed areas of high ink accumulation, generated by the occasional harder pressure of the pen on the paper at resting and corner points, respectively. In order to maximize spectral response, the laser was focused on one of these high ink concentration areas for SERS analysis. SERS spectra were taken in situ, by applying the Ag nanoparticle/PDMS substrate on the drawing with the Ag nanoparticle facing the analytical surface followed by direct back laser illumination. The SERS spectrum of Cabanis' red drawing is shown in Fig. 5b and was characterized

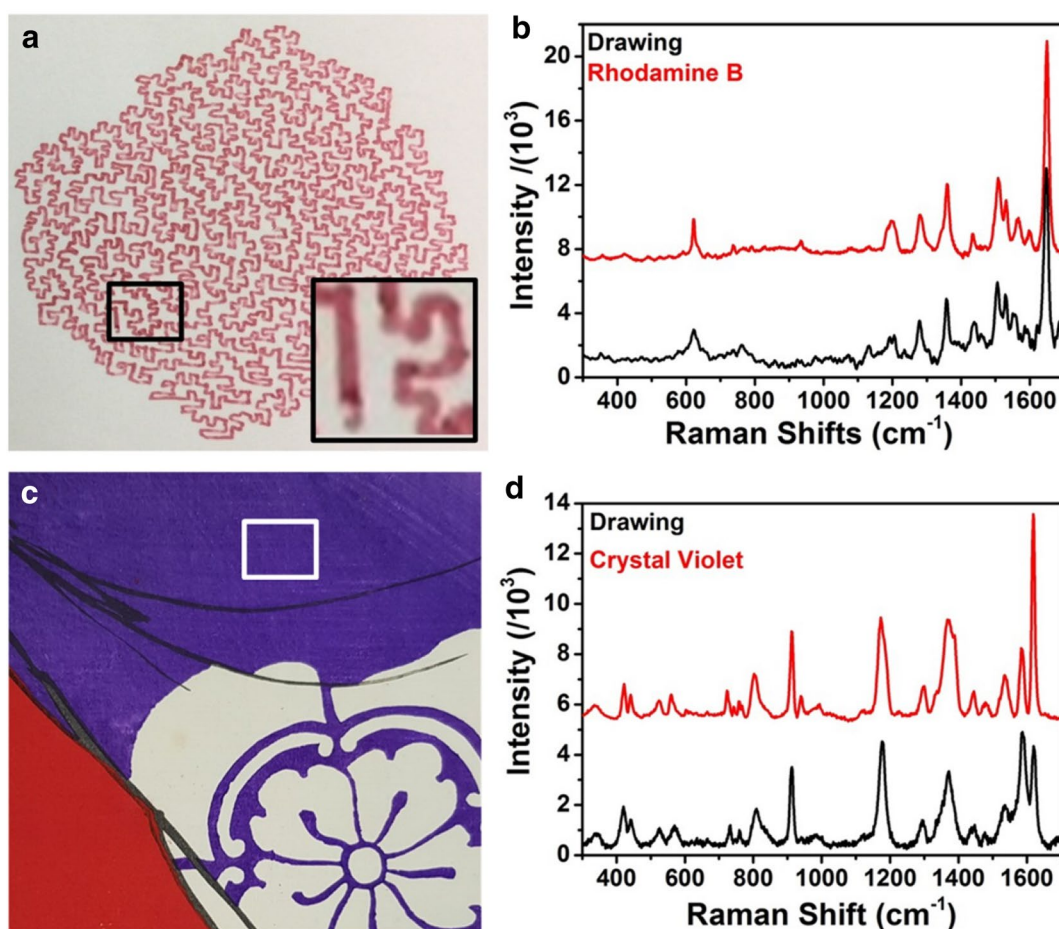


Fig. 5 **a, c** Photographs of red BIC pen drawing and Japanese wood print analyzed by Ag nanoparticle/PDMS composites. Insets: areas probed by SERS; **b** SERS spectra of the red ink drawing showed in **a** and reference dye Rodhamine B); **d** SERS spectra of the purple area showed in **c** and reference dye Crystal violet

by one strong band at 1648 cm^{-1} and medium intensity bands at $1532, 1505, 1429, 1352, 1282, 1198$ and 623 cm^{-1} all attributed to Rodhamine B. Figure 5c shows a photograph of a Japanese wood print characterized by areas of strong purple and red colorations. Figure 5d show SERS spectrum of the Japanese print marked purple area recorded by application of Ag nanoparticle/PDMS substrate followed by in situ back laser illumination. The SERS spectrum showed clear bands attributable to Crystal violet at 1622 and 1587 cm^{-1} (stretching of benzene rings), 1171 and 912 cm^{-1} (asymmetric stretching and bending of $\text{C}-\text{C}_{\text{center}}-\text{C}$ bonds), 442 and 419 cm^{-1} (bending of the $\text{C}-\text{N}-\text{N}$ bonds). Rodhamine B and Crystal violet reference spectra were also added to Fig. 5b, d, respectively.

Conclusions

Non-invasive SERS substrates were fabricated by transfer of self-assembled Ag nanoparticles on transparent substrates (glass and PDMS). The deposition of such SERS substrates on analytical surfaces allowed in situ acquisition of SERS spectra by laser back illumination directly on the investigated object without any need for micro-extraction steps. In contrast to featureless Raman spectra, high intensity and diagnostic SERS spectra were collected for drawings made by red, green and purple BIC ball-point pens. Following tests ensuring the minimal invasiveness of developed probes, flexible and transparent Ag nanoparticle/PDMS composites were applied to real works of art and SERS spectra obtained for red and purple dyes. The use of SERS has been limited to the necessity for micro-extraction steps, due to the impossibility to deposit SERS probes directly on the analytical surface. As well as transparency, Ag nanoparticle/PDMS composites displayed good adhesion to analytical surfaces and were

used for the in situ analysis of a modern art drawing and a Japanese woodblock print on paper. Compatibility with both bench and handheld Raman instrumentations was showed, suggesting the potential for Ag nanoparticle/PDMS composites to be used in the future for analysis of irregular surfaces as well as direct application to flat vertical surfaces. These results further promote the applicability of SERS as highly diagnostic tool for conservation science. The high transparency, conformability and flexibility of developed SERS probes allows their deposition on arbitrary and irregular surfaces thus substantially widening the range of SERS applications for analysis and preservation of art works.

Supplementary information

Supplementary information accompanies this paper at <https://doi.org/10.1186/s40494-019-0332-4>.

Additional file 1. Characterisation of Ag nanoparticles, UV-vis spectra of BIC pens and reference dyes, Blank SERS spectra, SERS spectrum of 4-aminobenzenethiol.

Abbreviations

4-ABT: 4-aminobenzene thiol; CuPh: copper phthalocyanine; CV: crystal violet; LDI-TOF-MS: Laser Desorption/Ionization Time of Flight Mass Spectrometry; MeOH: methanol; ODA: octadecylamine; PDMS: poly(dimethylsiloxane); SEM: scanning electron microscope; SERS: Surface Enhanced Raman Scattering; EASI-MS: Easy Ambient Sonic-spray Ionization Mass Spectrometry.

Authors' contributions

AA has performed all experimental work; AM has provided the artwork and supported spectral interpretation; DI has written the manuscript. All authors read and approved the final manuscript.

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Availability of data and materials

All data generated or analyzed during this study are included in this published article and its supplementary information files.

Competing interests

The authors declare that they have no competing interests.

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